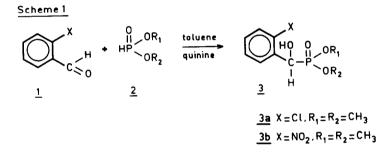
THE ABSOLUTE CONFIGURATION OF α -HYDROXYPHOSPHONATES

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<u>Summary</u>: The absolute configuration and CD spectra of a series of chiral α -hydroxyphosphonates and phosphoric acids have been determinated. The hydroxyphosphonates were prepared via a quinine catalyzed 1,2-addition reaction of phosphite to aldehyde.

Recently we described a new example of a quinine catalyzed asymmetric synthesis, namely the condensation between aromatic aldehydes and phosphites.¹



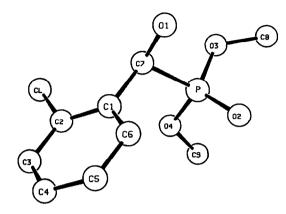
With $X=NO_2$ and $R_1=R_2 = t$ -butyl, ee's of 80-85% were achieved. Furthermore conversion of <u>3</u> to its mono-acid ($R_1=H$, $R_2=Me$) or diacid ($R_1=R_2=H$) readily furnished a series of chiral acids, useful in resolutions.²

Finally diesters of 3 ($R_1 = R_2 = Me$ or i-Pr) were found to be useful reagents for derivatization of chiral acids.³¹P-NMR the enantiomeric excess of chiral acids could be determinated.

Here we describe the determination of the absolute configuration of products resulting from the reaction of <u>1</u> with <u>2</u>. When X=Cl, the resulting α -hydroxyphosphonates are obtained optically active with ee's ranging from 10% (when $R_1=R_2=Me$) to 21% (when $R_1=R_2=i-Pr$). This trend towards higher ee's upon increasing the size of phosphite reactant appears to be general. Purification of <u>3a</u> proved to be exceedingly simple, only one recrystallization of material (ee = 10% as checked via both the ¹⁹F- and ³¹P-NMR using the ester prepared with Moscher reagent³) from acetone is necessary to obtain <u>3a</u> having ee's of better than 98%.

This unusual facility of purification by entrainment⁴ is undoubtelly due to the fact that <u>3a</u> as well as <u>3b</u> exist as conglomerates.⁵ Crystals of optically pure <u>3a</u> $[\alpha]_{78}^{20} = -74$ (c = 1, CHCl₃) were used for the determination of the absolute configuration using Bijvoet's method.⁶ Using Cu-radiation, we find that <u>3a</u> $[\alpha]_{578}^{20} = -74$ (c = 1, CHCl₃) has the absolute

configuration S(-) as shown in figure 1.

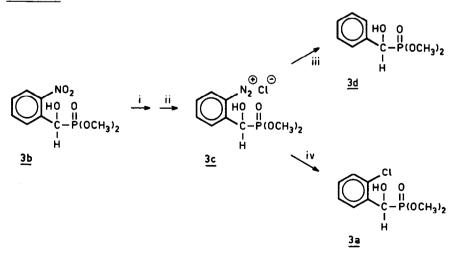


Relevant bond distances (A) and bond angles (degrees)⁷. C(1)-C(7) 1.515(7); C(7)-O(1) 1.437(6); C(7)-P 1.825(5); P-O(2) 1.459(4); P-O(3) 1.578(5); P-O(4) 1.579(4); C(6)-C(1)-C(7) 120.0(5); C(1)-C(7)-O(1) 111.7(4); C(1)-C(7)-P 105.3(3); C(7)-P-O(2) 116.5(3); C(7)-P-O(4) 101.6(2); C(7)-P-O(3) 103.3(3); O(3)-P-O(4) 105.0(3); O(4)-P-O(2) 113.8(3); O(2)-P-O(3) 115.0(3).

Figure 1 S(-) <u>3a</u>

Correlation between dimethyl o-chloro- α -hydroxybenzylphosphonate $\underline{3a} \left[\alpha\right]_{578}^{20}$ -74 (c = 1, CHCl₃) having the S(-) configuration and other α -hydroxyphosphonates was made using the following series of reactions.⁸

Scheme 2



 $\underline{3b} \left[\alpha\right]_{578}^{20} = -417 \text{ (c = 1, CHCl}_{3}\text{), m.p. } 108.5^{\circ}-110^{\circ}\text{C}^{1}\text{; } \underline{3d} \left[\alpha\right]_{578}^{20} = -46 \text{ (c = 1, acetone), m.p.} \\ 100^{\circ}-101^{\circ}\text{C}\text{; } \underline{3a} \left[\alpha\right]_{578}^{20} = -76 \text{ (c = 1, CHCl}_{3}\text{) m.p. } 114^{\circ}-115^{\circ}\text{; i. } 3 \text{ atm. H}_{2}\text{, } 5\% \text{ Pd/C, MeOH, } 3 \text{ hrs,} \\ \text{r.t.; ii. NaNO}_{2}\text{, HCl/H}_{2}\text{O, } 0^{\circ}\text{C}\text{; iii. } 20\% \text{ H}_{3}\text{PO}_{2}\text{, } 24 \text{ hrs, r.t.; iv. CuCl, } 3 \text{ hrs; } \underline{3b} \rightarrow \underline{3a} \text{ c.y. } 85\% \text{.} \\ \underline{3b} \rightarrow \underline{3a} \text{ c.y. } 86\% \text{.} \\ \underline{3b} \rightarrow \underline{3a} \text{ c.y. } 86\% \text{.} \\ \underline{3b} \rightarrow \underline{3a} \text{ c.y. } 86\% \text{.} \\ \underline{3b} \rightarrow \underline{3a} \text{ c.y. } 86\% \text{.} \\ \underline{3b} \rightarrow \underline{3a} \text{ c.y. } 86\% \text{.} \\ \underline{3b} \rightarrow \underline{3a} \text{ c.y. } 86\% \text{.} \\ \underline{3b} \rightarrow \underline{3a} \text{ c.y. } 86\% \text{.} \\ \underline{3b} \rightarrow \underline{3a} \text{ c.y. } 86\% \text{.} \\ \underline{3b} \rightarrow \underline{3a} \text{ c.y. } 86\% \text{.} \\ \underline{3b} \rightarrow \underline{3a} \text{ c.y. } 86\% \text{.} \\ \underline{3b} \rightarrow \underline{3a} \text{ c.y. } 86\% \text{.} \\ \underline{3b} \rightarrow \underline{3a} \text{ c.y. } 86\% \text{.} \\ \underline{3b} \rightarrow \underline{3a} \text{ c.y. } 86\% \text{.} \\ \underline{3b} \rightarrow \underline{3b} \text{ c.y. } 86\% \text{.} \\ \underline{3b} \rightarrow \underline{3b} \text{ c.y. } 86\% \text{.} \\ \underline{3b} \rightarrow \underline{3b} \text{ c.y. } 86\% \text{.} \\ \underline{3b} \rightarrow \underline{3b} \text{ c.y. } 86\% \text{.} \\ \underline{3b} \rightarrow \underline{3b} \text{ c.y. } 86\% \text{.} \\ \underline{3b} \rightarrow \underline{3b} \text{ c.y. } 86\% \text{.} \\ \underline{3b} \rightarrow \underline{3b} \text{ c.y. } 86\% \text{.} \\ \underline{3b} \rightarrow \underline{3b} \text{ c.y. } 86\% \text{.} \\ \underline{3b} \rightarrow \underline{3b} \text{ c.y. } 86\% \text{.} \\ \underline{3b} \rightarrow \underline{3b} \text{ c.y. } 86\% \text{.} \\ \underline{3b} \rightarrow \underline{3b} \text{ c.y. } 86\% \text{.} \\ \underline{3b} \rightarrow \underline{3b} \text{ c.y. } 86\% \text{.} \\ \underline{3b} \rightarrow \underline{3b} \text{ c.y. } 86\% \text{.} \\ \underline{3b} \rightarrow \underline{3b} \text{ c.y. } 86\% \text{.} \\ \underline{3b} \rightarrow \underline{3b} \text{ c.y. } 86\% \text{.} \\ \underline{3b} \rightarrow \underline{3b} \text{ c.y. } 86\% \text{.} \\ \underline{3b} \rightarrow \underline{3b} \text{ c.y. } 86\% \text{.} \\ \underline{3b} \rightarrow \underline{3b} \text{ c.y. } 86\% \text{.} \\ \underline{3b} \rightarrow \underline{3b} \text{ c.y. } 86\% \text{.} \\ \underline{3b} \rightarrow \underline{3b} \text{ c.y. } 86\% \text{.} \\ \underline{3b} \rightarrow \underline{3b} \text{ c.y. } 86\% \text{.} \\ \underline{3b} \rightarrow \underline{3b} \text{ c.y. } 86\% \text{.} \\ \underline{3b} \rightarrow \underline{3b} \text{ c.y. } 86\% \text{.} \\ \underline{3b} \rightarrow \underline{3b} \text{ c.y. } 86\% \text{.} \\ \underline{3b} \rightarrow \underline{3b} \text{ c.y. } 85\% \text{.} \\ \underline{3b} \rightarrow \underline{3b} \text{ c.y. } 85\% \text{.} \\ \underline{3b} \rightarrow \underline{3b} \text{ c.y. } 85\% \text{.} \\ \underline{3b} \rightarrow \underline{3b} \text{ c.y. } 85\% \text{.} \\ \underline{3b} \rightarrow \underline{3b} \text{ c.y. } 85\% \text{.} \\ \underline{3b} \rightarrow \underline{3b} \text{ c.y. } 85\% \text{.} \\ \underline{3b} \rightarrow \underline{3b} \text{ c.y. } 85\% \text{.} \\ \underline{3b} \rightarrow \underline{3b} \text{ c.y.$

Optically pure <u>3b</u> is converted with 5% Pd/C and 3 atm. H₂ to the less stable amino compound <u>3</u> $(X=NH_2, R_1=R_2=Me) [\alpha]_{578}^{20} = +49$ (c = 1, CHCl₃) which could be converted with NaNO₂ to the diazonium salt <u>3c</u>. Treatment of <u>3c</u> with a freshly prepared copper (I) chloride solution gives optically pure <u>3a</u> which was shown to be identical with the optically pure compound we obtained after the recrystallization of the condensation product of o-chlorobenzaldehyde and dimethylphosphite using quinine as a catalyst.

Treatment of diazonium salt <u>3c</u> with a 20% hypophosphoric acid solution leads to the optically pure compound <u>3d</u>. We prepared <u>3d</u> also by resolution of the mono-acid <u>3</u> $(X=R_1=H, R_2=CH_3)^2$ with quinine. The mono-acid <u>3</u> $(X=R_1=H, R_2=CH_3)$ is converted to the diester <u>3d</u> with diazomethane.⁹ This diester was identical with the compound we obtained by the diazonium salt route (scheme 2).

This data indicates that:

- (a) No racemization (at the potentially sensitive benzyl carbon in $\underline{3}$)¹⁰ has taken place during the conversion of the -NO₂ group to the other groups.
- (b) The absolute configuration of all of these compounds is now known based on the knowledge that $\underline{3a} \left[\alpha\right]_{578}^{20} = -74$ (c = 1, CHCl₃) has the S(-)-configuration.
- (c) Based on the CD spectra of $\underline{3}$ (X=H, NO₂, F, C1, Br or NH₂, R₁=R₂=Me) the α hydroxyphosphonates having the S configuration all show a negative Cotton effect at 225 nm. The CD spectrum of optically $\underline{3}c$ is shown in fig. 2.

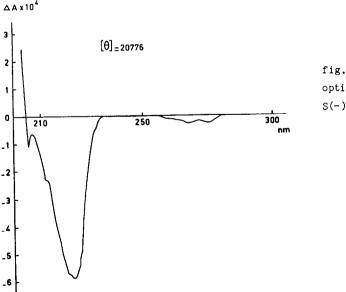


fig. 2: The CD spectrum of optically pure <u>3c</u>, having the S(-) configuration

Acknowledgment

This investigation was supported by the Netherlands Foundation for Chemical Research (SON) with financial aid from the Netherlands Organization for Advancement of Pure Research (ZWO).

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(Received in UK 22 October 1984)